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Inorganic Chemistry 2018

DOI: <https://doi.org/10.1021/acs.inorgchem.7b03175>

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DOI link to article:

<https://doi.org/10.1021/acs.inorgchem.7b03175>

Date deposited:

15/03/2018



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1,2,4-Triazines in the Synthesis of Bipyridine Bisphenolate ONNO Ligands and Their Highly Luminescent Tetradentate Pt(II) Complexes for Solution-Processable OLEDs

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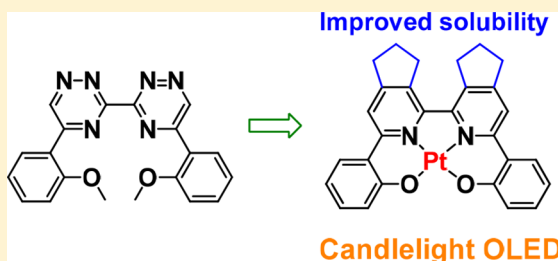
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Supporting Information

ABSTRACT: This article describes a convenient method for the synthesis of ONNO-type tetradentate 6,6'-bis(2-phenoxy)-2,2'-bipyridine (bipyridine bisphenolate, BpyBph) ligands and their platinum(II) complexes. The methodology includes the synthesis of 1,2,4-triazine precursors followed by their transformation to functionalized pyridines by the Boger reaction. Two complementary routes employing 3,3'- and 5,5'-bis-triazines allow a modification of the central pyridine rings in different positions, which was exemplified by the introduction of cyclopentene rings. The new ligands were used to prepare highly luminescent ONNO-type Pt(II) complexes. The position of the cyclopentene rings significantly influences the solubility and photophysical properties of these complexes. Derivatives with closely positioned cyclopentene rings are soluble in organic solvents and proved to be the best candidate for solution-processable organic light-emitting devices (OLEDs), showing efficient single-dopant candlelight electroluminescence.



INTRODUCTION

Tetradentate dianionic salen ONNO ligands are well known in coordination chemistry, and their metal complexes have found applications in a variety of scientific fields ranging from catalysis to bioinorganic and materials chemistry. For example, highly luminescent Pt(II) complexes of salen have been used as dopants in organic light-emitting diodes (OLEDs).¹ The tetradentate nature of the ligand reduces the possibility of ligand-exchange reactions and benefits both the stability and the luminescence properties of these complexes. Classical salen ligands are easily prepared by the condensation of salicylic aldehyde derivatives with the corresponding diamines. However, the resulting C=N imine bonds are prone to hydrolysis and other degradation pathways. To solve this problem, 6,6'-bis(2-phenoxy)-2,2'-bipyridine (bipyridine bisphenolate, BpyBph) were developed as a more robust system. Indeed, ONNO-type Pt(II) complexes of BpyBph and their phenanthroline analogues (PhenBph) showed strong luminescence in solution at room-temperature, good performance in OLEDs, and excellent thermal stability² (Figure 1). It should be noted that the performance of the complexes strongly depends on the rigidity and the substitution pattern of the ligand, with values of the photoluminescent quantum yield in solution varying from as low as 5% in BpyBph systems to as high as 58% in PhenBph-based complexes. Clearly, for BpyBph platinum(II) complexes there is room for the improvement of luminescence properties.

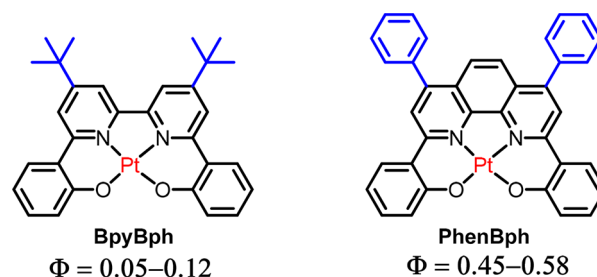
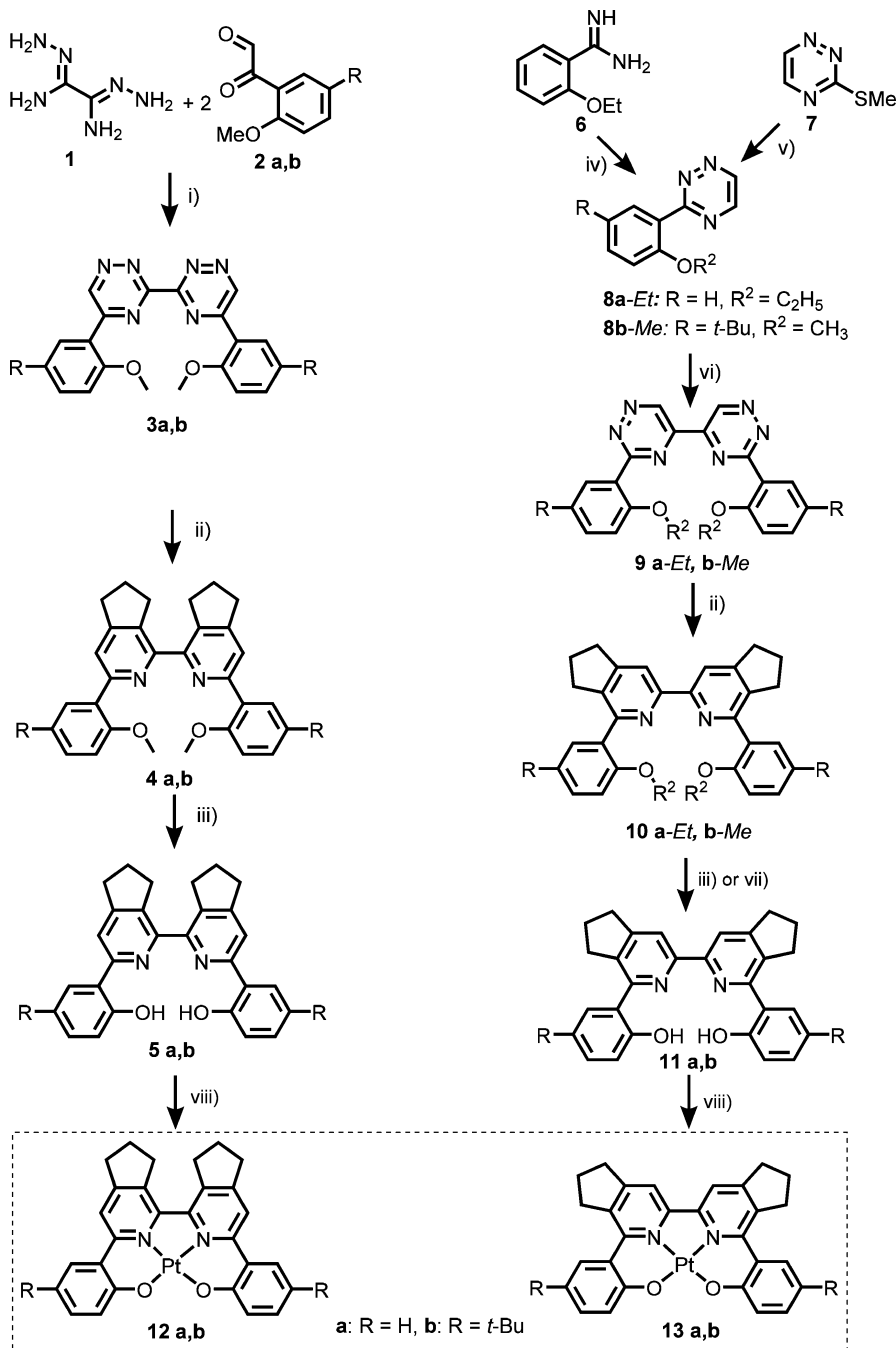


Figure 1. Structures of bipyridine bisphenolate (BpyBph) and phenanthroline bisphenolate (PhenBph) ligands reported by Che and co-workers.²

The synthesis of BpyBph ligands is more challenging than that of salen ligands, and there are only a few examples of such ligands described in the literature. The synthesis involves the reaction of 2,2'-bipyridine with 2-methoxyphenyllithium to give 6,6'-bis(2-methoxyphenyl)-2,2'-bipyridine, which is then demethylated to give the target ligand.³ Because highly nucleophilic reagents are used (*n*-BuLi), only a limited number of functional groups can be tolerated in the ligand. In this article, we demonstrate a convenient method for the synthesis of BpyBph ligands based on the chemistry of 1,2,4-triazines. This methodology was

Received: December 18, 2017

Scheme 1. Synthesis of Salen-Type BpyBph Proligands and Their Pt(II) Complexes^a

^aReaction conditions: (i) ethanol, reflux, 4–24 h; (ii) either 1-morpholino- or 1-pyrrolidinocyclopentene, please refer to experimental part for exact reaction conditions for a particular isomer; (iii) pyridinium chloride, 200–210 °C, 2–6 h; (iv) hydrazine hydrate, ethanol, 0–25 °C for 14 h then 40% glyoxal, reflux 1 h; (v) 3-*t*-butyl-5-methoxyphenyl boronic acid, CuMeSal, Na₂CO₃, Pd(PPh₃)₄, THF, reflux; (vi) KCN, 1,4-dioxane/water, RT, 14 h; (vii) BBr₃, RT, 14 h; (viii) K₂PtCl₄, AcOH, reflux, 14–16 h.

previously employed to prepare a variety of polypyridine ligands⁴ as well as their cyclometalated analogues.⁵ The key aspect of the methodology is that 1,2,4-triazines can be easily prepared from readily available starting materials and then transformed to their corresponding pyridines by the Boger reaction with enamines. In this way, a variety of substituents can be introduced into the pyridine ring. In this article, we report two synthesis routes employing 3,3'- and 5,5'-bis-triazines as key building blocks in preparing four novel ligands. We then show that the ligands readily form Pt(II) complexes, which are highly luminescent and

can successfully be used as dopants in solution-processed OLEDs. In this study, we concentrate exclusively on solution-processed OLEDs and demonstrate that the position of cyclopenteno units plays a key role in increasing the solubility of the complexes.

RESULTS AND DISCUSSION

Synthesis. The synthesis of the novel salen-type proligands is depicted in Scheme 1. Because the N1 and N2 atoms of the triazine ring are lost during the Boger reaction as dinitrogen, we are limited to using 3,3'- and 5,5'-bis-triazines as key building

blocks in the synthesis. 3,3'-Bis-triazines **3a** and **3b** were prepared in moderate yields by the reaction of corresponding glyoxal derivatives **2a** and **2b** with ethanedihydrazonamide (**1**). In order to prepare 5,5'-bis-triazine (**9a-Et**), we started with commercially available 2-ethoxybenzamidine (**6**), which was treated with 1 equiv of hydrazine hydrate to give the corresponding amidrazone.^{6,7} The amidrazone was not isolated but was used in situ in the reaction with glyoxal to give **8a-Et** in 66% yield. *tert*-Butyl derivative **8b-Me** was synthesized by the desulfurative Suzuki reaction of sulfide **7** with commercially available 3-*tert*-butyl-5-methoxyphenyl boronic acid.⁸ The next step involved the dimerization of 1,2,4-triazines **8a-Et** and **8b-Me** by the action of potassium cyanide. This high-yielding reaction was first reported in 1973⁹ and has already been applied to the synthesis of bipyridines.¹⁰ 3,3'-Bis-triazines **3a** and **3b** and 5,5'-bis-triazines **9a-Et** and **9b-Me** were then converted to 2,2'-bipyridines **4a** and **4b** and **10a-Et** and **10b-Me**, respectively, by the Boger reaction with 1-morpholinocyclopentene. Finally, *O*-demethylation with pyridinium chloride or boron tribromide under standard conditions gave the desired salen-type pro-ligands **5a** and **5b**, and **11a** and **11b**. Complexes **12a** and **12b**, and **13a** and **13b** were then prepared by the reaction of the corresponding proligands with potassium tetrachloroplatinate in refluxing acetic acid.

It was immediately apparent that **12a** is far more soluble than **13a**. One explanation for this might be that the cyclopentene units in **12a** cause a slight twist of the pyridine rings, reducing the planarity of the molecule and thus minimizing the intermolecular interaction. Indeed, computational data supports this argument and shows that **12a** is less planar than **13a** (Figure 2).

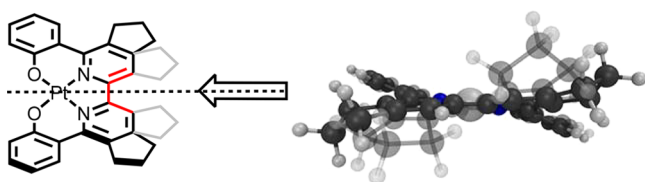


Figure 2. Comparison of optimized structures **12a** (gray) and **13a** (black).

Interestingly, the introduction of the *tert*-butyl groups did not significantly improve the solubility in both series of complexes in toluene. Only **12a** dissolves well in toluene, while other complexes have very limited solubility. For this reason, all of the solution-processed OLEDs in this work were made using chlorobenzene as the solvent.

Luminescent Properties. The small changes in the ligand structure between **12** and **13** cause noticeable differences in the photophysical properties of the complexes. For example, in toluene solution, **13a**, which is an isomer of **12a** differing only by the position of the cyclopentene ring, has a more blue-shifted emission spectrum relative to that of **12a**. A similar trend is observed in **12b** and **13b**.

The photoluminescence spectra in solution (Figure 3) are nearly featureless and relatively broad, but when dispersed in a rigid and nonpolar polymer (Zeonex), the spectra show a more resolved vibronic structure. Interestingly, the phosphorescence spectra of each pair of analogues in Zeonex (**12a** and **12b**, **13a** and **13b**) are nearly identical, although they differ in solution. This indicates that the difference in the emission color of *tert*-butylated and nonsubstituted analogues is mostly a result of additional vibrational/rotational modes in *tert*-butylated molecules and the electron-donating effect of this group is less important.

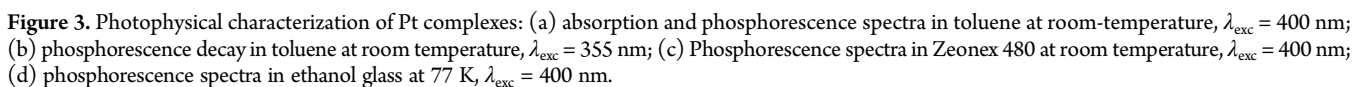
At 77 K, all four complexes display well-resolved, blue-shifted emission. In all cases, the 0–0 transition is the strongest in the spectrum. Within each group of complexes (**12a** and **12b**, **13a** and **13b**), the low-temperature phosphorescence spectra show a very similar shape with identical vibronic structure, but the analogues with *tert*-butyl groups (**12b**, **13b**) give more red-shifted spectra than do their counterparts (**12a** and **13a**). This is, in fact, an indication that the electronic effect of the *tert*-butyl group is significant. In general, notable rigidochromism is observed in these complexes, showing that the molecules are susceptible to intramolecular motion and interactions at room temperature. In both series, the introduction of the *tert*-butyl group reduces the luminescence yield, Φ_{PL} , which is probably due to the larger number of vibrational/rotational modes. Molecules with cyclopentene rings on the outer periphery of the bipyridine moiety (**13a**, **13b**) have a larger Φ_{PL} than do the corresponding isomers (**12a**, **12b**).

This argument is supported by the phosphorescence lifetime measurements. In toluene solution, **13a** shows the longest phosphorescence lifetime, whereas **12b** shows the shortest. Shorter phosphorescence time constants in **12b** and **13b** are also found compared to their corresponding analogues (**12a** and **13a**). We attribute the shorter phosphorescence lifetime of **12b** and **13b** relative to their analogues (**12a** and **13a**) to the increased nonradiative decay. All decay profiles recorded in toluene are perfectly monoexponential as shown in Figure 3. The photophysical properties of complexes **12a**, **12b**, **13a**, and **13b** are summarized in Table 1.

Electrochemistry. Cyclic voltammetry studies on complexes **12a**, **12b**, **13a**, and **13b** show that all compounds behave similarly and show a reversible reduction that involves the NN (bipyridine) ligands. The oxidation of the complexes is, however, irreversible. In this case, the behavior of **12a** and **13a** is nearly identical, and this is also observed for their *tert*-butylated analogues, **12b** and **13b** (Figure 4, Table 2). This indicates that the oxidation of these complexes involves the phenyl rings of the OO ligand, as alkyl substitution of this moiety changes the electrooxidation signal pattern. This means that the alkyl- and nonalkyl-substituted complexes undergo different chemical reactions upon oxidation. The complexes have very similar oxidation and reduction potentials, but **13a** and **13b** have a larger electrochemical energy gap (0.05–0.07 eV), which explains the observed blue shift of their emission in comparison to that of the other two molecules.

Theory and Computational Chemistry. To support the interpretation of the experimental results, we have performed density functional theory (DFT) and time-dependent density functional theory (TDDFT) simulations on **12a** and **13a**, with the primary focus of rationalizing the structural and emission properties. The structure of the lowest triplet states was optimized using DFT(B3LYP)¹² and a def2-SVP basis set as implemented within the ORCA quantum chemistry package.¹³ The excited-state energies were computed using linear response-time-dependent density functional theory (LR-TDDFT) using the Tamm–Damcoff approximation and the B3LYP exchange and correlation functional as implemented within the Amsterdam Density Functional (ADF) code.¹⁴ The SOC matrix elements were computed with the perturbative approach developed by Wang and Ziegler.¹⁵ A DZP basis set was used for all atoms, except Pt (TZP), and scalar relativistic effects were accounted for using the zeroth-order relativistic approximation (ZORA).¹⁶

Figure 5 shows HOMO and LUMO orbitals, which constitute the dominant transitions responsible for both the S_1 and T_1 excited states. These correspond to charge-transfer states, from



solvent	absorption					emission				
	$\lambda_{\text{abs}}, \text{nm}$ ($\varepsilon \times 10^{-4}, \text{M}^{-1} \text{cm}^{-1}$) ^a					$\lambda_{\text{em}}, \text{nm}$ ^b	$\tau, \mu\text{s}$ ^c	Φ_{PL} ^d	$\lambda_{\text{em}}, \text{nm}$ ^e	$\tau, \mu\text{s}$ ^f
	toluene					toluene			Zeonex 480	
12a	541 (0.22), 509 (0.31), 478 (0.27), 417 (0.80), 329 (1.85), 308 (2.12)					624	1.76 ± 0.08	0.24	580	7.6 ± 0.4
13a	527 (0.17), 489 (0.27), 399 (0.85), 320 (1.80)					593	2.98 ± 0.14	0.34	564	5.5 ± 0.4
12b	564 (0.30), 523 (0.38), 488 (0.30), 436 (1.02), 335 (2.20), 310 (2.61)					646	0.93 ± 0.04	0.12	589	5.4 ± 0.2
13b	547 (0.15), 506 (0.26), 470 (0.30), 410 (0.80), 323 (1.57)					610	1.42 ± 0.18	0.25	572	4.8 ± 0.4

and dopant. (see SI for more details about fabrication of devices.) The dopant concentration was limited in some cases due to the poor solubility of some of the complexes in chlorobenzene. For this reason, toluene could also not be used as a solvent, except for **12a**, which was dissolved in toluene to produce a device with a 5% dopant concentration. Devices of **12a** in toluene had performances that were similar to those of the presented chlorobenzene-based OLEDs. The poor solubility of complexes **13a** and **13b** resulted in a limited dopant concentration (1 and 3%, respectively). This is likely to decrease the efficiency of host to dopant energy transfer (due to a limited Forster and Dexter radii) and cause problems with direct charge injection into the dopant, which in turn decreases the efficiency of the device. This is clearly

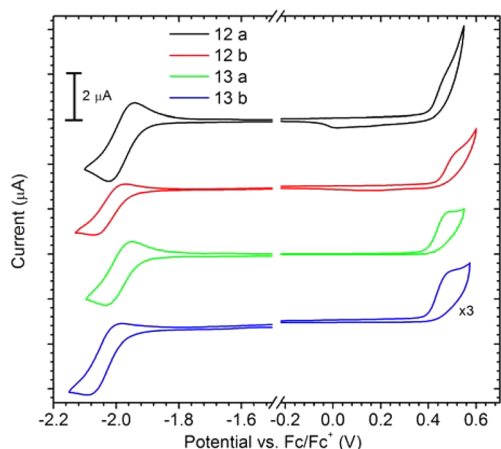


Figure 4. Electrochemical oxidation and reduction cycles of the investigated complexes recorded with cyclic voltammetry at a 0.05 V s^{-1} scan rate in $0.1 \text{ M Bu}_4\text{NBF}_4/\text{DCM}$ supporting electrolyte. Note that the results for **13b** were recorded at a lower concentration due to its limited solubility.

Table 2. Electrochemical Properties of the Investigated Complexes^a

complex	$E_{\text{Ox}}^{\text{CV}}, \text{V}^b$	$E_{\text{Red}}^{\text{CV}}, \text{V}^c$	IP, eV^d	EA, eV^e	$E_g^{\text{el}}, \text{eV}^f$
12a	0.40	−1.92	5.50	3.18	2.32
13a	0.44	−1.95	5.54	3.15	2.39
12b	0.40	−1.92	5.50	3.18	2.32
13b	0.40	−1.97	5.50	3.13	2.37

^aNote All Potentials Are Relative to the Standard Potential of Ferrocene/Ferrocenium Redox Couple. ^bOnset oxidation potential was recorded with cyclic voltammetry. ^cOnset reduction potential was recorded with cyclic voltammetry. ^dIonization potential $\text{IP} = E_{\text{ox}}^{\text{CV}} + 5.1$. ^eElectron affinity $\text{EA} = E_{\text{red}}^{\text{CV}} + 5.1$. ^fElectrochemical energy gap $E_g^{\text{el}} = \text{IP} - \text{EA}$. IP and EA correspond to the HOMO and LUMO energy, respectively, and are used as their estimates.

the problem in devices containing **13a**. This is also shown in the photophysical study where, in the case of PVK:PBD films with low doping concentrations (complexes **13a** and **13b**), another, slower mechanism for energy transfer is observed (i.e., limited by hole–electron pair recombination¹⁷). When the photophysical data is compared with electroluminescence spectra it becomes clear that in OLEDs direct charge injection to the dopant is the dominant process, while energy transfer from host to guest is less dominant. On the other hand, energy transfer governs the photoluminescence in a PVK:PBD mixed host. For this reason, it is important that the efficiency of the OLED devices is slightly larger than it would be expected to be from photophysics.

The compounds perform well as dopants in OLEDs (Figure 6), giving electroluminescence from yellow-green to orange that is in full agreement with the photophysical results. All dopants except **12b** give efficient devices with an EQE of 8.4–10.4% and a satisfactory current efficiency of $22.3\text{--}24.9 \text{ cd A}^{-1}$. The brightness of these devices reached 7000 cd m^{-2} in the case of **12a**, while the device derived from **12b** shows the lowest brightness of 2500 cd m^{-2} . The turn-on voltage is in the typical range for the OLED structure used with a PVK:PBD host.¹⁸

Because of the higher photoluminescence quantum yield both in doped film and in solution, the **13a**-based OLED was expected to have a higher maximum EQE and in general to outperform the **12a**-based OLED. However, the low solubility of the complex imposed limitations in doping concentration, and the device with **13a** as the dopant (1%) is therefore less efficient than the device incorporating **12a** (5%). The yellowish electroluminescence of devices 1 and 3 is close to that of candlelight, while device 2 has a more green and device 4 a more red electroluminescence. The colorimetric coordinates in the CIE 1931 color space of a typical candle ($0.52, 0.42$)¹⁹ and those of presented devices 1 and 3 are close to this value (see Table 3). Interestingly, the candlelight color has been achieved with the use of only one emitter whereas other solutions use 2 to 3 emitters.^{19–21}

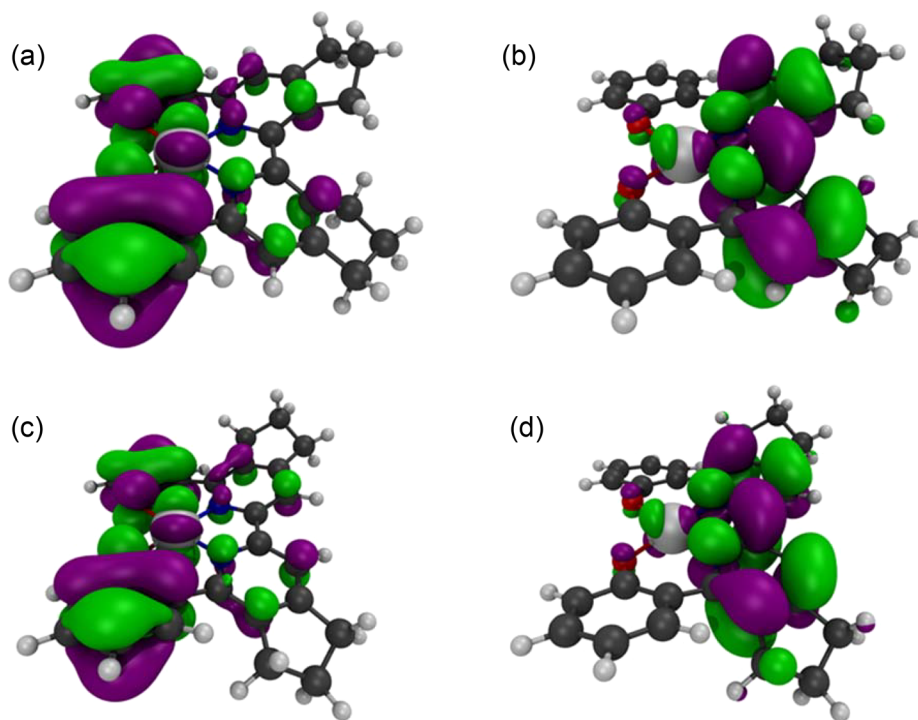


Figure 5. HOMO (a and c) and LUMO (b and d) orbitals of **12a** (upper) and **13a** (lower).

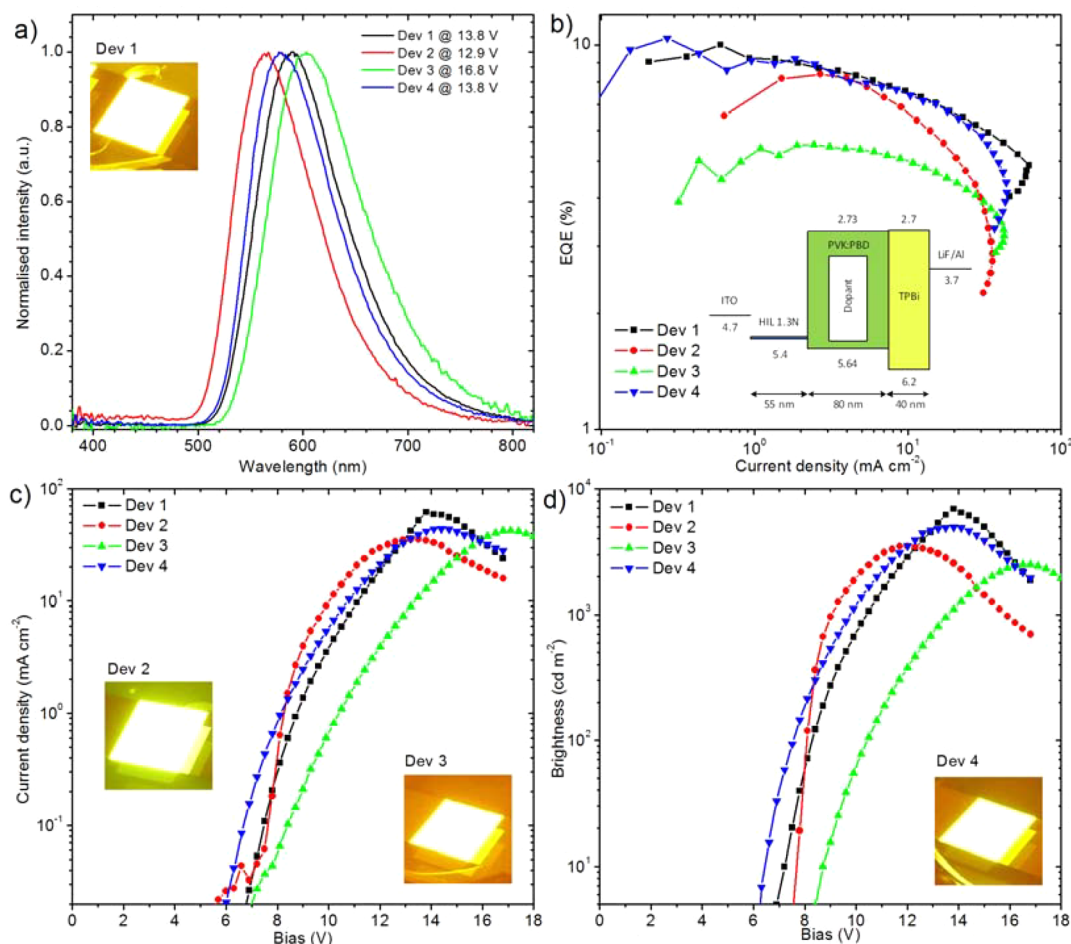


Figure 6. Characteristics of the OLED devices produced from **12a**, **12b**, **13a**, and **13b**: (a) electroluminescence spectra; (b) external quantum efficiency (EQE) vs current density; (c) J - V curves; and (d) brightness vs bias. The insets are photographs of the working devices. The inset in panel (b) shows the device architecture.

Table 3. Summary of the Characteristics of the OLED Devices^a

device	dopant	turn on at 10 cd m ⁻² , V	EQE, %			max. current efficiency, cd A ⁻¹	max. brightness, cd m ⁻²	EL maximum, nm	CIE (x, y) at max brightness
			max.	at 100 cd m ⁻²	at 1000 cd m ⁻²				
Dev 1	12 a	7.2	10.0	9.7	8.2	23.2	7000	590	0.54, 0.46
Dev 2	13 a	7.6	8.4	6.6 ^b	8.2	24.9	3500	567	0.46, 0.51
Dev 3	12 b	8.7	5.5	5.3	4.9	9.9	2500	603	0.57, 0.42
Dev 4	13 b	6.4	10.4	9.3	8.0	22.3	5000	577	0.52, 0.47

^aDevice architecture: ITO|HIL 1.3N (55 nm)|dopant (1–5%) co PVK:PBD (60:40) (80 nm)|TPBi (40 nm)|LiF (0.8 nm)|Al (100 nm). Doping concentration: Dev 1: **12a** (5%), Dev 2: **13a** (1%), Dev 3: **12b** (5%), Dev 4: **13b** (3%). ^bThe value was recorded at 118 cd m⁻².

CONCLUSIONS

The synthesis methodology based on the chemistry of 1,2,4-triazines proved to be efficient and versatile in the synthesis of tetradentate dianionic bipyridine bisphenolate ONNO ligands. The scope of the methodology goes beyond luminescent complexes and is likely to attract attention from scientists in other fields. Many other functional groups can be introduced by the Boger reaction, opening up additional pathways for optimizing and tuning ONNO ligands for different applications. In Pt(II) complexes formed by the new ligands, the introduction of the *tert*-butyl groups was found not to improve solubility in toluene but reduced the quantum efficiency of the complexes and is thus not a good strategy for optimizing the dopants. However, the 1,2,4-triazine methodology illustrated herein does allow for the

synthesis of more soluble isomers by positioning the cyclopentene rings closer together, leading to less-planar structures. This significantly improves the solubility of the complex, but the detrimental effects on the photophysical properties are less profound. This allowed us to obtain the required doping concentration to achieve efficient single-dopant candlelight OLED devices.

EXPERIMENTAL SECTION

The description of the materials and methods and experimental as well as experimental procedures for *tert*-butyl analogues are described in the Supporting Information.

5,5'-Bis(2-methoxyphenyl)-3,3'-bi-1,2,4-triazine (3a). (1*Z*,2*Z*)-Ethanedihydrazonamide (**1**) (1.16 g, 10 mmol) was suspended in ethanol (75 mL), and a 1 M solution of 2-methoxyphenylglyoxal (**2a**) in ethanol (22 mL, 22 mmol) was added. The reaction mixture was

heated under reflux for 4 h. The mixture was allowed to cool to ambient temperature, and the solid was filtered off, washed with ethanol (20 mL) followed by diethyl ether (40 mL), and allowed to dry in air to afford title compound **3a** as a yellow-green solid (1.87 g, 50%). ^1H NMR (400.1 MHz, CDCl_3 , Me_4Si) δ : 10.12 (s, 2H), 8.38 (d, J = 7.6, 2H), 7.55 (t, J = 7.6, 2H), 7.16 (t, J = 7.6, 2H), 7.06 (d, J = 8.8, 2H), 3.99 (s, 6H); ^{13}C NMR (100.6 MHz, CDCl_3 , Me_4Si) δ : 161.9, 158.7, 155.8, 150.1, 134.1, 132.2, 122.3, 121.8, 111.6, 55.8. HRMS (ESI) calcd for $\text{C}_{20}\text{H}_{17}\text{N}_6\text{O}_2$ ($[\text{M} + \text{H}]^+$): m/z 373.1413. Found: m/z 373.1411.

6,6'-Bis(2-methoxyphenyl)-2,2'-bipyridine (4a). 5,5'-Bis(2-methoxyphenyl)-3,3'-bi-1,2,4-triazine, **3a** (2.3 g, 6.18 mmol), was placed in a round-bottomed flask and 1-morpholinocyclopentene (3.78 g, 4 equiv) was added. The flask was heated at 140 °C for 3 h and was then allowed to cool to 70 °C. Methanol (30 mL) was added and the mixture was stirred at reflux for 10 min and then allowed to cool to room-temperature. The solid was filtered and washed with MeOH (30 mL) to give desired product **4a** (2.2 g, 79%). ^1H NMR (400.1 MHz, CDCl_3 , Me_4Si) δ : 7.82 (dd, J = 7.6, 1.6, 2H), 7.67 (s, 2H), 7.30–7.35 (m, 2H), 7.05 (t, J = 7.2, 2H), 6.98 (d, J = 7.6, 2H), 3.84 (s, 6H), 3.33 (t, J = 7.6, 4H), 2.99 (t, J = 7.6, 4H), 2.11 (quin, J = 7.6, 4H); ^{13}C NMR (100.6 MHz, CDCl_3 , Me_4Si) δ : 157.1, 152.6, 153.5, 155.0, 138.0, 131.5, 130.1, 129.3, 120.9, 120.5, 111.5, 55.7, 33.0, 32.4, 25.2. HRMS (ESI) calcd for $\text{C}_{30}\text{H}_{29}\text{N}_2\text{O}_2$ ($[\text{M} + \text{H}]^+$): m/z 449.2229. Found: m/z 449.2233.

6,6'-Bis(2-hydroxyphenyl)-2,2'-bipyridine (5a). A mixture of **4a** (1.80 g, 0.4 mmol) and pyridinium chloride (18.0 g, 155.7 mmol) was stirred at 210 °C for 6 h. The mixture was allowed to cool to room temperature, and water (50 mL) was added. The solid was filtered off and washed sequentially with water and methanol to give **5a** as a solid. Yield: 1.5 g, (90%). ^1H NMR (400.1 MHz, CDCl_3 , Me_4Si) δ : 7.89 (s, 2H), 7.86 (dd, J = 8.4, 1.2, 2H), 7.29 (t, J = 8.4, 2H), 6.99 (dd, J = 8.4, 0.8, 2H), 6.92 (t, J = 8.4, 2H), 3.13 (t, J = 7.6, 4H), 3.08 (t, J = 7.6, 4H), 2.14 (quin, J = 7.6, 4H); ^{13}C NMR (100.6 MHz, CDCl_3 , Me_4Si) δ : 159.7, 158.0, 155.0, 148.4, 138.0, 131.1, 126.2, 119.2, 118.8, 118.4, 115.2, 33.4, 31.6, 25.5. HRMS (ESI) calcd for $\text{C}_{28}\text{H}_{23}\text{N}_2\text{O}_2$ ($[\text{M} + \text{H}]^+$): m/z 421.1916. Found: m/z 421.1902.

Platinum Complex 12a. A mixture of proligand **5a** (105 mg, 0.25 mmol) and potassium tetrachloroplatinate (104 mg, 0.25 mmol) in acetic acid (40 mL) was heated under reflux for 14 h. The solvent was evaporated. The product was purified by column chromatography using silica gel and a mixture of DCM/ethyl acetate (10/1). Yield 61 mg (40%). ^1H NMR (400.1 MHz, CDCl_3 , Me_4Si) δ : 7.65 (s, 2H), 7.61 (br. d, J = 8.3, 2H), 7.29 (m, 2H), 7.24 (dd, J = 8.4, 1.1, 2H), 6.69 (m, 2H), 3.08 (t, J = 7.1, 4H), 2.78 (t, J = 7.1, 4H), 2.22 (quin, J = 7.1, 4H). HRMS (TOF) calcd for $\text{C}_{28}\text{H}_{23}\text{N}_2\text{O}_2$ ^{194}Pt ($[\text{M} + \text{H}]^+$): m/z 613.1386. Found: m/z 613.1387.

3-(2-Ethoxyphenyl)-1,2,4-triazine (8a-Et). To an ice-cold, stirred suspension of 2-ethoxybenzamidino hydrochloride (10 g, 50 mmol) in ethanol (50 mL), hydrazine hydrate (2.5 g, 50 mmol) was added. The mixture was stirred at 0 to 15 °C overnight. A 40% aqueous glyoxal solution (8.7 g, 0.6 mol, 1.2 equiv) was added. The reaction was stirred for 20 min at room temperature and then gradually heated to reflux. The mixture was heated under reflux for 1 h. Water (250 mL) was added, and the mixture was extracted with DCM (3 \times 30 mL). The organic phase was dried over MgSO_4 , filtered, and concentrated to a volume of 20 mL. The product was purified by column chromatography on silica gel using DCM/ethyl acetate 1/1 as an eluent. The solvent was evaporated to give a product as a yellow oil. Yield 6.63 g (66%). Analytical characteristics are identical to literature data.^{8a}

3,3'-Bis(2-ethoxyphenyl)-5,5'-bi-1,2,4-triazine (9a-Et). 3-(2-Ethoxyphenyl)-1,2,4-triazine (**8a-Et**) (2.01 g, 10 mmol) was stirred in a mixture of water (50 mL) and dioxane (20 mL) until complete dissolution occurred. KCN (850 mg, 13 mmol) was added in two portions, 30 min apart. The mixture was stirred at rt overnight. The solid was filtered off and washed with water and methanol to give **9a-Et**. Yield 1.65 g (82%). ^1H NMR (400 MHz, CDCl_3) δ : 10.25 (s, 2H), 7.99 (dd, J = 7.6, 1.8, 2H), 7.54 (ddd, J = 8.3, 7.6, 1.8, 2H), 7.16 (ddd, J = 8.3, 8.3, 1.0, 2H), 7.12 (br. d, J = 8.3, 2H), 4.20 (q, J = 6.9, 4H), 1.43 (t, J = 6.9, 6H); ^{13}C NMR (100 MHz, CDCl_3) δ : 165.4, 157.9, 150.8, 143.96, 132.7, 132.1, 124.6, 120.8, 113.3, 64.5, 14.9. HRMS (ESI) calcd for $\text{C}_{22}\text{H}_{21}\text{N}_6\text{O}_2$ ($[\text{M} + \text{H}]^+$): m/z 401.1726. Found: m/z 401.1713.

6,6'-Bis(2-ethoxyphenyl)-2,2'-bipyridine (10a-Et). 5,5'-Bis(2-ethoxyphenyl)-3,3'-bi-1,2,4-triazine (**9a-Et**) (2.2 g, 5.5 mmol) was placed in a round-bottomed flask, and 1-pyrrolidinocyclopentene (7.54 g, 55 mmol, 10 equiv) was added. The flask was held at 140 °C for 15 min. Acetic acid (30 mL) was added, and the mixture was stirred at reflux for 10 min. Methanol (30 mL) and water (30 mL) were added, and the mixture was allowed to cool to room temperature. The solid was filtered and washed with MeOH (30 mL) to give the desired product (2.1 g, 80%). The product was used without further purification. However, the product can be purified by recrystallization from acetic acid/methanol mixture. ^1H NMR (400.1 MHz, CDCl_3 , Me_4Si) δ : 8.30 (s, 2H), 7.51 (dd, J = 7.6, 1.8, 2H), 7.35 (ddd, J = 8.3, 7.6, 1.8, 2H), 7.07 (br. t, J = 7.6, 2H), 7.12 (br. d, J = 7.6, 2H), 4.02 (q, J = 7.0, 4H), 2.98 (t, J = 7.4, 4H), 2.83 (t, J = 7.4, 4H), 2.04 (quin, J = 7.4, 4H), 1.28 (t, J = 6.9, 6H); ^{13}C NMR (100 MHz, CDCl_3) δ : 156.3, 154.97, 154.4, 152.2, 139.6, 131.7, 130.7, 129.4, 120.7, 116.3, 112.3, 64.0, 33.2, 31.5, 25.2, 15.0. HRMS (ESI) calcd for $\text{C}_{32}\text{H}_{33}\text{N}_2\text{O}_2$ ($[\text{M} + \text{H}]^+$): m/z 477.2542. Found: m/z 477.2533.

BpyBph proligand (11a). At room temperature, BBr_3 (4 mL) was slowly added dropwise to the solution of **10a-Et** (500 mg, 1.05 mmol) in CH_2Cl_2 (20 mL). After the addition was complete, the mixture was stirred at room temperature under argon for 18 h. The reaction was quenched by the slow addition of water. DCM (30 mL) was added, and the organic phase was separated, washed with water, dried, and filtered. The filtrate was filtered through a plug of silica gel. The silica gel was washed with DCM. The solvent was evaporated, and the residue was treated with methanol (5 mL) and filtered. Yield 190 mg (43%) ^1H NMR (400.1 MHz, CDCl_3 , Me_4Si) δ : 8.07 (s, 2H), 7.76 (dd, J = 7.8, 1.6, 2H), 7.33 (ddd, J = 8.2, 7.8, 1.6, 2H), 7.09 (dd, J = 8.2, 1.1, 2H), 6.95 (br. t, J = 8.2, 2H), 3.33 (t, J = 7.4, 4H), 3.09 (t, J = 7.4, 4H), 2.19 (quin, J = 7.4, 4H). HRMS (ESI) calcd for $\text{C}_{28}\text{H}_{25}\text{N}_2\text{O}_2$ ($[\text{M} + \text{H}]^+$): m/z 421.1916. Found: m/z 421.1899.

Platinum Complex (13a). A mixture of the ligand (126 mg, 0.3 mmol) and potassium tetrachloroplatinate (124 mg, 0.3 mmol) in acetic acid (30 mL) was heated under reflux for 14 h. Water (30 mL) was added, and the solid product was filtered off, washed with water, and dried to give **13a** as a yellow solid. Yield 160 mg (87%). ^1H NMR (400.1 MHz, CDCl_3 , Me_4Si) δ : 7.73 (s, 2H), 7.49 (dd, J = 8.1, 1.4, 2H), 7.41 (dd, J = 8.3, 1.1, 2H), 7.29 (m, 2H), 6.62 (m, 2H), 3.26 (t, J = 7.4, 4H), 2.75 (t, J = 7.4, 4H), 1.82 (qu, J = 7.4, 4H). HRMS (TOF) calcd for $\text{C}_{28}\text{H}_{23}\text{N}_2\text{O}_2$ ^{194}Pt ($[\text{M} + \text{H}]^+$): m/z 613.1386. Found: m/z 613.1395.

■ ASSOCIATED CONTENT

Supporting Information

The Supporting Information is available free of charge on the ACS Publications website at DOI: 10.1021/acs.inorgchem.7b03175.

Details of the synthesis of *tert*-butyl isomers **12b** and **13b**, copies of NMR and MASS spectra, additional details for photophysical and electrochemical characterization, and OLED device fabrication (PDF)

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Notes

The authors declare no competing financial interest.

■ ACKNOWLEDGMENTS

This work was supported by Northumbria University, EU Horizon 2020 research and innovation program under Marie

Sklodowska-Curie grant agreement no. 674990 (EXCILIGHT), and EPSRC, EP/L02621X/1, and EP/P004873/1. We thank the EPSRC UK National Mass Spectrometry Facility at Swansea University for recording the high-resolution mass spectra.

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- (21) In response to a reviewer's comments on the unusually high EQE and CE values based on an estimation from the PLQY, the measurements have been repeated and were found to show reproducible values. The detailed mechanism and reasons for these may require further study.